

# K<sub>p</sub>

Pre-lesson assignment – Textbook p298-301

## Define the following terms

- Mole Fraction
- Partial Pressure

## Make notes on using the equilibrium constant

*Use the following questions as guidance*

1. Write an equation for calculation of a mole fraction.
2. Write an equation for calculation of partial pressure.
3. Demonstrate how the partial pressure of N<sub>2</sub>, O<sub>2</sub> and other gases can be calculated in air.
4. Write an expression for K<sub>p</sub> for the reaction of hydrogen gas and iodine gas.
5. Explain why expressions of K<sub>p</sub> only include gases.