

Y12/13 – Equilibria

Stick this checklist into your yellow book at the beginning of the topic. Tick off the topics as you cover them.		
In this module you are expected to be able to...	In Notes	Revised
<ul style="list-style-type: none">Define the terms<ul style="list-style-type: none">Dynamic equilibriumHeterogeneous equilibriumHomogeneous equilibrium		
<ul style="list-style-type: none">Use Le Chatelier’s principle to explain how change in conditions leads to change in position of equilibrium		
<ul style="list-style-type: none">Explain why equilibrium position is unaffected by catalysts.		
<ul style="list-style-type: none">Explain how compromises are reached when determining optimum conditions for industrial processes.		
<ul style="list-style-type: none">Write an expression for K_c and K_p for homogeneous and heterogeneous equilibria		
<ul style="list-style-type: none">Use K_c to estimate the position of equilibrium.		
<ul style="list-style-type: none">Use the terms ‘mole fraction’ and ‘partial pressure’		
<ul style="list-style-type: none">Calculate quantities present at equilibrium using K_c or K_p given data, or calculate K_c or K_p given data		
<ul style="list-style-type: none">Describe the effect of temperature, concentration and pressure on K_p and K_c.		
<ul style="list-style-type: none">Explain why molar quantities in an equilibrium change when conditions are altered with reference to K_c and K_p.		

Pre-test Evaluation

I have...	
Updated my yellow book notes	
Ensured I understand all of my notes	
Looked on the open drive for additional work	
Asked my teacher for guidance	
Confidence rating	I’m doomed! -- - = + ++ I am the BOSS!