Buffers 2

Pre-lesson assignment: Textbook page 335-339

Make notes on buffer solutions

Use the following questions as guidance

1. Give a worked example of how the pH of a buffer can be calculated given K_a and the concentration of the acid and salt present.
2. Give a worked example of how the pH of a buffer can be calculated given K_a and the initial concentration and volume of both acid and salt solutions used to make it.
3. Give a worked example of how to calculate pH of a buffer made by part neutralisation
4. State the effects of pH of blood being outside the normal range.
5. State the normal range of pH of blood.
6. Explain how hydrogen carbonate ions buffer the blood.