

Y13 – pH and Buffers

Stick this checklist into your yellow book at the beginning of the topic. Tick off the topics as you cover them.		
In this module you are expected to be able to...	In Notes	Revised
<ul style="list-style-type: none">Define the terms<ul style="list-style-type: none">Acid and Base in terms of Brønsted-LowryMonobasic, dibasic and tribasicWeak/Strong in the context of an acid.		
<ul style="list-style-type: none">Identify acid-conjugate base and base-conjugate acid pairs.		
<ul style="list-style-type: none">Recall the reactions of acids with oxides, hydroxides, carbonates and metals, including ionic equations.		
<ul style="list-style-type: none">Calculate pH from<ul style="list-style-type: none">[H⁺] from a strong acidK_w and [OH⁻] from a strong baseK_a and [HA] from a weak acidK_a [HA] and [A⁻] from a buffer solution		
<ul style="list-style-type: none">Calculate pK_a		
<ul style="list-style-type: none">Understand how K_a is related to the strength of a weak acid.		
<ul style="list-style-type: none">Explain the action of a buffer solution in terms of equilibrium, when [H⁺] is increased and decreased, and how this causes pH to remain more or less constant over a small range of pH		
<ul style="list-style-type: none">Draw pH titration curves for a combination of strong and weak acids and bases.		
<ul style="list-style-type: none">Explain why indicators change colour in terms of equilibrium.		
<ul style="list-style-type: none">Select suitable indicators for an acid-base titration.		

Pre-test Evaluation

I have...	
Updated my yellow book notes	
Ensured I understand all of my notes	
Looked on the open drive for additional work	
Asked my teacher for guidance	
Confidence rating	I’m doomed! -- - = + ++ I am the BOSS!