

Redox Titrations – $\text{S}_2\text{O}_3^{2-}$

Pre-lesson assignment – Textbook p. 381-385

Make notes on redox reactions involving thiosulfate ions

1. Write an equation for the reduction of iodine with thiosulfate ions, showing both half equations.
2. Write a procedure for the titration of sodium thiosulfate against an oxidising agent. Include:
 - a. Why potassium iodide is added.
 - b. An explanation of why starch is added near the end point.
 - c. The colour change at the End point.
3. Read through the green box and worked example, and check that you can follow the stoichiometry.
4. Explain how brass or copper samples are treated in order to produce iodine, and how this can give a percentage of copper in brass or impure copper